

Acids and Alkalis

	<u>Working towards Mastery (W)</u>	<u>Meeting Mastery (M)</u>	<u>Beyond Mastery (B)</u>
Acids and Alkalis	<ul style="list-style-type: none">- The pH of a solution depends on the strength of the acid: strong acids have lower pH values than weak acids.- Mixing an acid and alkali produces a chemical reaction, neutralisation, forming a chemical called a salt and water.- Acids have a pH below 7, neutral solutions have a pH of 7, alkalis have a pH above 7.- Acids and alkalis can be corrosive or irritant and require safe handling.- Hydrochloric, sulfuric and nitric acid are strong acids. Acetic and citric acid are weak acids.	<ul style="list-style-type: none">- Identify the best indicator to distinguish between solutions of different pH, using data provided.- Use data and observations to determine the pH of a solution and explain what this shows.- Explain how neutralisation reactions are used in a range of situations.- Describe a method for how to make a neutral solution from an acid and alkali.	<ul style="list-style-type: none">- Given the names of an acid and an alkali, work out the name of the salt produced when they react.- Deduce the hazards of different alkalis and acids using data about their concentration and pH.- Estimate the pH of an acid based on information from reactions.